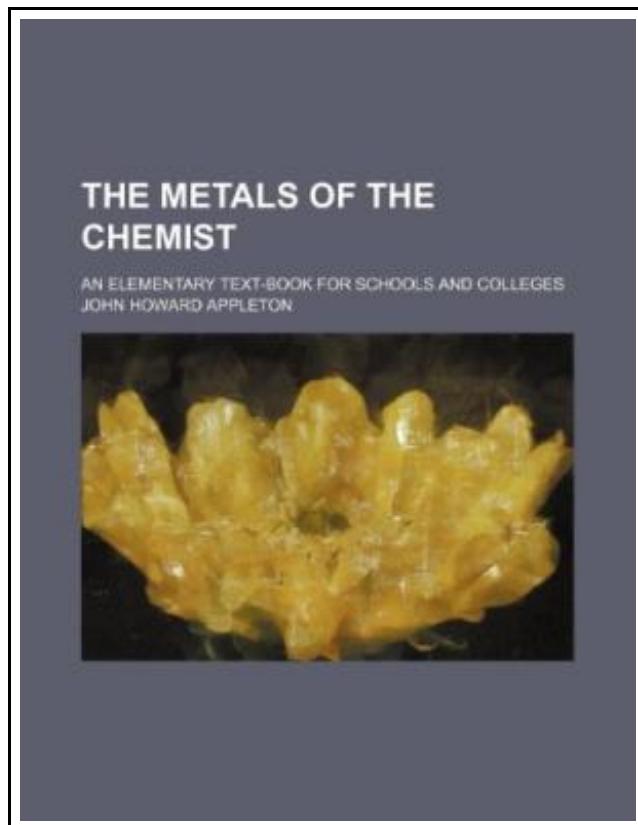


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1899 Excerpt: .ridges with a view of exposing a larger surface to the action of the gas. After the lime has been properly spread in these chambers, chlorine gas is allowed to enter in sufficient quantity to fill them. It is then left to stand in contact with the lime for two or three days. As a result a considerable portion of the gas is absorbed. If, upon now testing some of the powder, it is found not yet fully saturated with gas, chlorine is again supplied to the chamber and allowed to remain there for another period of two or three days. The lime is usually then found to have absorbed so much chlorine as will make up somewhere from thirty to forty per cent, of the bleaching-powder produced. Fig. 62. Longitudinal vertical section of English chlorine still, made of slabs of stone. Manganese dioxide, and hydrochloric acid (introduced at g) liberate chlorine (which is conducted away by m). While, as already intimated, the exact composition of bleaching powder has not yet been conclusively made out, it is sufficient for all ordinary purposes to consider it as a combination of three compounds which may be expressed thus: $\text{CaCl}_2\text{O}_2 + \text{CaCl}_2 + \text{CaO}_2\text{H}_2$ Calcium hypochlorite Calcium chloride Calcium hydroxide It must be admitted to be the fact that all ordinary chemical changes in which bleaching-powder takes part, may be expressed without consideration of the calcium hydroxide given in the formula. This latter substance, however, appears to be always...

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